

Antwoorden opstellen reactievergelijkingen 4H + 4V

- $4 \text{Fe(s)} + 3 \text{O}_2\text{(g)} \rightarrow 2 \text{Fe}_2\text{O}_3\text{(s)}$
- $2 \text{N}_2\text{(g)} + 3 \text{O}_2\text{(g)} \rightarrow 2 \text{N}_2\text{O}_3\text{(g)}$
- $2 \text{PCl}_5\text{(s)} \rightarrow 2 \text{P(s)} + 5 \text{Cl}_2\text{(g)}$
- $2 \text{Hg(l)} + \text{I}_2\text{(s)} \rightarrow 2 \text{HgI(s)}$
- $\text{SnS}_2\text{(s)} \rightarrow \text{Sn(s)} + 2 \text{S(s)}$
- $2 \text{AgBr(s)} \rightarrow 2 \text{Ag(s)} + \text{Br}_2\text{(g)}$
- $2 \text{Pb(s)} + \text{O}_2\text{(g)} \rightarrow 2 \text{PbO(s)}$
- $4 \text{HCl(g)} + \text{O}_2\text{(g)} \rightarrow 2 \text{H}_2\text{O(l)} + 2 \text{Cl}_2\text{(g)}$
- $\text{Fe}_2\text{O}_3\text{(s)} + 3 \text{H}_2\text{(g)} \rightarrow 2 \text{Fe(s)} + 3 \text{H}_2\text{O(l)}$
- $4 \text{Fe(s)} + 3 \text{O}_2\text{(g)} \rightarrow 2 \text{Fe}_2\text{O}_3\text{(s)}$
- $2 \text{C}_2\text{H}_6\text{(g)} + 7 \text{O}_2\text{(g)} \rightarrow 4 \text{CO}_2\text{(g)} + 6 \text{H}_2\text{O(l)}$
- $\text{PCl}_5\text{(s)} + 4 \text{H}_2\text{O(l)} \rightarrow \text{H}_3\text{PO}_4\text{(s)} + 5 \text{HCl(g)}$
- $\text{Na}_2\text{O(s)} + \text{H}_2\text{O(l)} \rightarrow 2 \text{NaOH(s)}$
- $\text{P}_2\text{O}_5\text{(s)} + 3 \text{H}_2\text{O(l)} \rightarrow 2 \text{H}_3\text{PO}_4\text{(s)}$
- $\text{N}_2\text{(g)} + 3 \text{H}_2\text{(g)} \rightarrow 2 \text{NH}_3\text{(g)}$
- $4 \text{Al(s)} + 3 \text{PbO}_2\text{(s)} \rightarrow 2 \text{Al}_2\text{O}_3\text{(s)} + 3 \text{Pb(s)}$
- $2 \text{KClO}_3\text{(s)} \rightarrow 2 \text{KCl(s)} + 3 \text{O}_2\text{(g)}$
- $\text{NH}_3\text{(g)} + \text{HCl(g)} \rightarrow \text{NH}_4\text{Cl(s)}$
- $\text{Na}_2\text{CO}_3\text{(s)} + 2 \text{HCl(g)} \rightarrow 2 \text{NaCl(s)} + \text{H}_2\text{O(l)} + \text{CO}_2\text{(g)}$
- $4 \text{NH}_3\text{(g)} + 5 \text{O}_2\text{(g)} \rightarrow 4 \text{NO(g)} + 6 \text{H}_2\text{O(l)}$
- $2 \text{Fe(s)} + 3 \text{S(s)} \rightarrow \text{Fe}_2\text{S}_3\text{(s)}$
- $\text{Fe(s)} + \text{S(s)} \rightarrow \text{FeS(s)}$
- $\text{Cr}_2\text{O}_3\text{(s)} + 3 \text{Zn(s)} \rightarrow 2 \text{Cr(s)} + 3 \text{ZnO(s)}$
- $2 \text{C}_{12}\text{H}_{22}\text{O}_{11}\text{(s)} \rightarrow 24 \text{C(s)} + 22 \text{H}_2\text{(g)} + 11 \text{O}_2\text{(g)}$
 - $\text{C}_{12}\text{H}_{22}\text{O}_{11}\text{(s)} \rightarrow 12 \text{C(s)} + 11 \text{H}_2\text{O(l)}$
- $\text{C}_{12}\text{H}_{22}\text{O}_{11}\text{(s)} + 12 \text{O}_2\text{(g)} \rightarrow 12 \text{CO}_2\text{(g)} + 11 \text{H}_2\text{O(l)}$